

Quiz 3: Periodic Properties
(10 pts)

1. What happens (generally) to Z_{eff} as you go down the periodic table? (2 pt)

stays the same

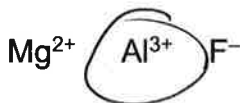
2. What happens to size as you go across the periodic table? (2 pt)

decreases

Why does that happen? (2 pt)

b/c Z_{eff} increases

3. Circle the smallest ion (2 pt)



Why? (2 pt)

isoelectronic but Al^{3+} has most p^+