

In Class Assignment 8

Periodic Properties

Explain the following statement as clearly and as thoroughly as you can. **Use radial probability diagrams, electron configurations, etc. to support your answers whenever possible.**

1. Lithium has a higher ionization energy than sodium but a lower ionization energy than beryllium.
2. Titanium is smaller than scandium and calcium is smaller than potassium. However, there is a smaller size **difference** between titanium and scandium than between calcium and potassium.
3. The size of Cd^{2+} is smaller than that of Nb^{2+} but it is also smaller than Sn^{2+}